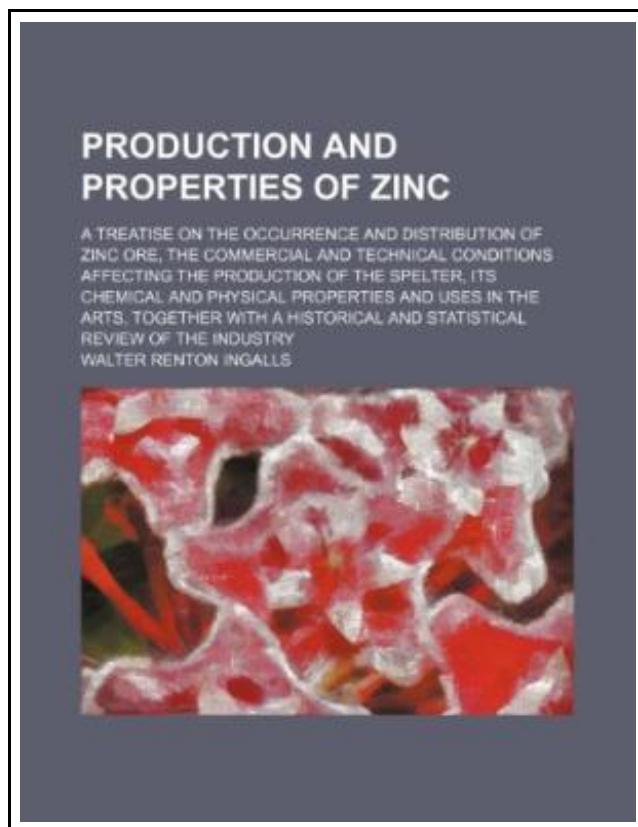


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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1902 Excerpt: . disengagement of carbonic acid and the precipitate formed is a basic hydrocarbonate. This hydrocarbonate is soluble in an aqueous solution of carbonic acid; upon exposure to the air a granular powder deposits from the solution, which has been taken for neutral carbonate of zinc by some chemists, but by others has been held to be a basic carbonate. According to Thorpe² neutral zinc carbonate, $ZnCO_3$, can be produced artificially by adding acid carbonate of sodium to a solution of zinc sulphate, while Roscoe and Schorlemmer state that it is produced by precipitation from $ZnSO_4$ with an excess of hydrogen potassium car 1 In a paper on Sulphide Ore Treatment, Metallurgy, London, June 19, 1901. read before the Institution of Mining and 2 Dictionary of Applied Chemistry, p. 1057. bonate ($KHCO_3$), the neutral carbonate of sodium (Na_2CO_3) throwing down a basic carbonate of zinc, which is the more basic the higher the temperature of precipitation and the more dilute the solution. Zinc carbonate, $ZnCO_3$, is produced when the solution of a zinc salt is precipitated by an alkaline carbonate under pressure. As obtained by heating a solution of $ZnCl_2$ to 150-160 C. and precipitating with calcium carbonate or sodium bicarbonate it is a white microcrystalline powder. The neutral hydrated carbonate ($2ZnCO_3 \cdot H_2O$) is an amorphous powder obtained by digesting a basic carbonate of zinc with ammonium bicarbonate. The anhydrous carbonate of zinc is soluble in carbonic acid water. A solution of CO_2 at five atmospheres, taking...

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